Unit D – Compounds Compared to Mixtures

Name

Per

Papers worth reviewing

- What Makes a Good Data Table?
- What makes a Good Graph?
- LAD D1 Getting to Know the Bunsen Burner
- LAD D2 Law of Constant Composition, SG
- LAD D3 Percent of Oxygen in Air
- LAD D4 Separating Mixtures
- LAD D5 Law of Constant Composition, MgO
- NoteSheet D1 Classifying Matter
- NoteSheet D2 Mass Ratios and Percent Composition
- Practice D1 Particulate Diagrams
- Practice D2 Working with Mass Ratios and Percent Composition
- Practice D3 Practice with LAD D5 Calcs
- Consider using the class presentation and clicker questions for review as well (available on the unit D document page at the top)
- Consider reviewing your openers. (You can find a pdf with all the openers and answers on the documents page)

Objectives

- 1. Recognize that matter can be classified as a pure substance or a mixture
 - pure substances are either elements or compounds
 - be able to describe both heterogeneous and homogeneous mixtures (aka solutions)
 - compare and contrast mixtures and compounds
 - · provide examples of various types of mixtures as well as compounds and elements
- 2. Know how to use a Bunsen burner
 - proper method of lighting and turning off the burner
 - adjusting the gas and adjusting the air
 - being able to identify a quality flame and what parts of the flame are hottest
- 3. Be able to state the Law of Constant Composition
- 4. Remembering the general procedures of the labs so that you can explain error
- 5. Recognizing and identifying measurements vs calculations when in the lab
- 6. Know how to perform all of the lab calculations
 - percent yield
 - percent error
 - mass ratios
- 7. Draw particulate diagrams describing
 - atoms
 - molecules (diatomic molecules)
 - chemical reactions with reactant & product quantities
- 8. Balance simple chemical equations
- 9. Working calculations involving mass ratios and percent composition

Unit D – Compounds Compared to Mixtures

Vocabulary List

- Law of Constant Composition (aka Law of Definite Proportions)
- heating to a constant mass
- experimental
- theoretical
- percent error
- percent yield
- pure substance
 - element
 - compound
- mixture
 - heterogeneous
 - homogeneous
- mass ratio
- percent composition (by mass)

This vocabulary list is meant to complement your study. Knowing this list alone, without the concepts on the front would not prepare you for the test.